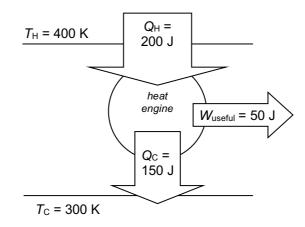
1. a)
$$\eta_{\text{Carnot}} = \frac{W_{\text{useful}}}{Q_{\text{H}}} = \frac{40 \text{ J}}{100 \text{ J}} = 0.40 = \underline{40 \%}$$

b)
$$\eta_{\text{Carnot}} = \frac{Q_{\text{H}} - Q_{\text{C}}}{Q_{\text{H}}} = \frac{100 \text{ J} - 60 \text{ J}}{100 \text{ J}} = 0.40 = \underline{40 \%}$$

c)
$$\eta_{\text{Carnot}} = \frac{T_{\text{H}} - T_{\text{C}}}{T_{\text{H}}} = \frac{500 \text{ K} - 300 \text{ K}}{500 \text{ K}} = 0.40 = \underline{40 \%}$$

2. a), b) and d)



b)
$$Q_C = Q_H - W_{useful} = 200 J - 50 J = 150 J$$

c)
$$\eta_{\text{Carnot}} = \frac{W_{\text{useful}}}{Q_{\text{H}}} = \frac{50 \text{ J}}{200 \text{ J}} = 0.25 = \underline{25 \%}$$

d)
$$\eta_{Carnot} = \frac{T_{H} - T_{C}}{T_{H}}$$
 $\eta_{Carnot} \cdot T_{H} = T_{H} - T_{C}$

$$T_{\rm C} = T_{\rm H} - \eta_{\rm Carnot} \cdot T_{\rm H} = T_{\rm H} \cdot (1 - \eta_{\rm Carnot}) = 400 \text{ K} \cdot (1 - 0.25) = 300 \text{ K}$$

3. The number above the fraction bar needs to be large. That is, for $\eta_{\text{Carnot}} = \frac{T_{\text{H}} - T_{\text{C}}}{T_{\text{H}}}$ the difference in temperature $(T_{\text{H}} - T_{\text{C}})$ ought to be as large as possible.

4. a)
$$\eta_{\text{Carnot}} = \frac{T_{\text{H}} - T_{\text{C}}}{T_{\text{H}}} = \frac{773 \text{ K} - 293 \text{ K}}{773 \text{ K}} = 0.621 = \underline{62.1 \%}$$

b)
$$W_{\text{useful}} = \eta_{\text{Carnot}} \cdot Q_{\text{H}} = 0.62 \cdot 2'600 \text{ kJ} = 1'614 \text{ kJ} = 1.61 \text{ MJ}$$

c)
$$Q_C = Q_H - W_{useful} = 2.6 \text{ MJ} - 1.6 \text{ MJ} = 986 \text{ kJ}$$

5. a)
$$\eta_{\text{Carnot}} = \frac{W_{\text{useful}}}{Q_{\text{H}}} = \frac{W_{\text{useful}}}{Q_{\text{C}} + W_{\text{useful}}}$$
 $\eta_{\text{Carnot}} \cdot (Q_{\text{C}} + W_{\text{useful}}) = W_{\text{useful}}$

 $\eta_{\text{Carnot}} \cdot Q_{\text{C}} + \eta_{\text{Carnot}} \cdot W_{\text{useful}} = W_{\text{useful}}$

$$\eta_{\text{Carnot}} \cdot Q_{\text{C}} = W_{\text{useful}} - \eta_{\text{Carnot}} \cdot W_{\text{useful}} = W_{\text{useful}} \cdot (1 - \eta_{\text{Carnot}})$$

$$Q_{\rm C} = \frac{W_{\rm useful} \cdot (1 - \eta_{\rm Carnot})}{\eta_{\rm Carnot}} = \frac{1.00 \text{ GJ} \cdot (1 - 0.30)}{0.30} = \underline{2.33 \text{ GJ}}$$

b)
$$T_C = T_H \cdot (1 - \eta_{Carnot}) = 473 \text{ K} \cdot (1 - 0.30) = 331 \text{ K}$$
 $\mathcal{G}_C = 58 \text{ °C}$

c) increase \mathcal{G}_{C} :

$$T_{\rm C} = T_{\rm H} \cdot (1 - \eta_{\rm Carnot}) = 473 \text{ K} \cdot (1 - 0.60) = 189 \text{ K}$$
 $g_{\rm C} = -84 \, {}^{\circ}\text{C}$

decrease g_H :

$$T_{\rm H} = \frac{T_{\rm C}}{1 - \eta_{\rm Carnot}} = \frac{331 \text{ K}}{1 - 0.60} = 828 \text{ K}$$
 $g_{\rm H} = \underline{555 \, ^{\circ}\text{C}}$